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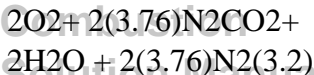
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Determine the
molecular weight and
stoichiometric mole and
mass air-fuel ratios for
the Oklahoma gas mole
composition given in
Table 3.4. Solution.

Equation (3.2), CH_4+

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shows that there are $2 + 2(3.76) = 9.52$ moles of air required for complete combustion of each mole of methane.

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